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Aggregation Behavior of Mono-endcapped Hydrophobically Modified Poly(sodium acrylate)s in Aqueous Solution

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Introduction

Water-soluble polymers find numerous applications in industry, since they can strongly modify the rheological properties of aqueous solutions. Handles for fine-tuning these properties include the molecular mass, charge, and chemical structure of the polymeric backbone. An important class of macromolecules are the hydrophobically modified polymers. In past years extensive research has been focused on random hydrophobically modified poly(sodium acrylate)s in aqueous solution. Polymers with molecular weights varying between 800 and 31,700 were synthesized by radical polymerization using an initiator and chain transfer agent. The resulting polymers form hydrophilic micromodains in aqueous solutions. The following conditions were applied: no salt and pH 5 and 9, respectively; 1 M sodium citrate and pH 9. At pH 5 the critical aggregation concentration (CAC, the concentration at which micromodains are formed) increases with increasing molecular weight of the polymers. The concentration range for aggregation is about 0.2-2.4 mM. At pH 9 the carboxylic acid groups are deprotonated and electrostatic repulsions are introduced; therefore the concentration for aggregation rises to about 80 mM. Interestingly, in case of polymers having \( M_n > 1400 \) the CAC decreases with increasing molecular weight due to a counterion-concentration gradient toward the hydrophobic micromodain. Near the micromodain the counterion binding is increased, reducing the electrostatic repulsions and allowing for lower aggregation concentrations. In the presence of 1 M sodium citrate this anomalous trend is suppressed to a large extent; since the overall counterion binding is increased and the CAC is lower. The concentration for aggregation is then in the same range as at pH 5 in the absence of salt.

Key Words: micelles; mono-endcapped hydrophobically modified poly(sodium acrylate); isothermal titration calorimetry; steady-state fluorescence.

EXPERIMENTAL

Materials. Mono-endcapped hydrophobically modified poly(sodium acrylate)s were prepared by radical polymerization using an initiator and chain transfer agent. The procedure has been described previously (19). The initiator, ditridecanoyl peroxide, was prepared from tridecanoic acid chloride using sodium hydroxide and hydrogen peroxide (20). Tridecanoic acid chloride was obtained by reacting tridecanoic acid with excess thionyl chloride (21). Tridecanoic acid was obtained from Aldrich and was used without further purification. The chain transfer agent, \( n \)-dodecylthiol, was obtained from Aldrich and was distilled before use. The monomer, acrylic acid, was obtained from Across and distilled before use. Due to the structure and properties of the initiator and chain transfer agent there are three possibilities for the linker between the hydrophilic tail and hydrophilic headgroup: a sulfur atom, an acid group, or no linker. For the...
low-molecular-weight polymers relatively much chain transfer agent has been used. Therefore the sulfur linker predominates. For the high-molecular-weight polymers only small amounts of chain transfer agent have been used. Therefore, there is mainly no linker. We assume that the acid linker is present only in small amounts. Non-hydrophobically modified poly(acrylate) was obtained from National Starch; the polymer was used without further purification. The carboxylic groups were converted into their sodium salts by increasing the pH to 9 with sodium hydroxide. The viscosity-averaged molecular mass was 8000 (22). Using the weight-averaged molecular mass and viscosity-averaged mass ratio of a similar polymer and assuming a polydispersity of 2, the number-averaged molecular mass was estimated to be 5100.

Pyrene (99% pure) was obtained from Aldrich and was used as received. It was dissolved in water and filtered at least 1 day after dissolution. This solution was then diluted once. No excimers were present, since steady-state fluorescence showed no peak near 450 nm, characteristic of pyrene excimers (23).

Sodium citrate was obtained from Merck and used as received. It was dissolved in water and filtered at least 1 day after dissolution. This solution was then diluted once. No excimers were present, since steady-state fluorescence showed no peak near 450 nm, characteristic of pyrene excimers (23).

Solutions were prepared using doubly distilled water. The pH was measured using an Orion Model SA 720 digital pH-meter. Stirred solutions of hydrophobically modified poly(sodium acrylate)s were titrated with a concentrated HCl solution (pH 0.23) at room temperature.

1H-NMR. 1H-NMR spectroscopy (Varian 300 MHz) was used to determine the number-averaged molecular weight of the mono-endcapped hydrophobically modified poly(sodium acrylate)s using the endgroup method (24). Recently this method has proven to be reliable (25). The results are shown in Table 1.

<table>
<thead>
<tr>
<th>Mw from</th>
<th>Mn from</th>
<th>Mw/Mn</th>
<th>Number of monomeric units (from NMR)</th>
</tr>
</thead>
<tbody>
<tr>
<td>NMR</td>
<td>GPC</td>
<td>GPC</td>
<td></td>
</tr>
<tr>
<td>790</td>
<td>6.6</td>
<td></td>
<td></td>
</tr>
<tr>
<td>1,490</td>
<td>1900</td>
<td>2700</td>
<td>1.4</td>
</tr>
<tr>
<td>3,020</td>
<td>3100</td>
<td>4300</td>
<td>1.4</td>
</tr>
<tr>
<td>6,720</td>
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<td></td>
<td>70</td>
</tr>
<tr>
<td>12,480</td>
<td></td>
<td></td>
<td>131</td>
</tr>
<tr>
<td>31,710</td>
<td></td>
<td></td>
<td>336</td>
</tr>
</tbody>
</table>

Steady-state fluorescence spectroscopy. Steady-state fluorescence spectroscopic measurements were performed using a SLM SPF-500C spectrofluorometer equipped with a thermostated cell holder and a magnetic stirring device. The instrument settings were as follows: excitation wavelength, 335 nm; slit width, 1 nm. The emission spectrum was recorded from 360 to 410 nm (slit width, 1 nm; step size, 0.20 nm; filter 2). Pyrene was used as a fluorescent probe at concentrations lower than 10⁻⁶ M.

Samples were prepared by dissolving the desired amount of polymer in aqueous solutions containing about 5 × 10⁻⁷ M pyrene. Solutions were allowed to equilibrate for at least 24 h. Then the emission spectrum was recorded and the intensities of the first (at about 370 nm) and third (at about 385 nm) peaks were determined. Control experiments were performed with solutions that had been equilibrated for 24, 48, and 72 h. All the experimentally determined I₁/I₃ ratios were the same within experimental error.

Measurements were performed under the following conditions: pH 4.90-5.10, no salt; pH 8.60-10, no salt; pH 8.60-10, 1 M sodium citrate. All experiments were done at 298 K.

Microcalorimetry (27). An Omega isothermal titration microcalorimeter (Microcal Inc., Northampton, MA) was used to measure the enthalpies of the polymer aggregation. The procedure was as follows. Thirty aliquots containing 10 μl of a polymer solution were injected into a sample cell initially containing solution S₀. The polymer interacts with S₀, and the resulting enthalpic effect is measured by comparing the differences in energy that have to be added to the measuring cell and to the reference cell (containing pure water) to keep both cells at a constant temperature. Concentrations were corrected for the added volumes. Microcal Origin software was used to process the data. All experiments were done at 298 K, unless stated otherwise.

**RESULTS AND DISCUSSION**

Potentiometric Titrations

Potentiometric titrations of weak polyacids are very complicated and protonation of and counterion binding to the acid groups depend on many parameters (e.g., molecular weight of the polymer, concentration, salt concentration, tacticity) (28). The degree of dissociation (α) of the hydrophobically modified mono-endcapped poly(sodium acrylate)s was calculated by using the difference between the amount of protons added and the number of protons free in the solution. At pH 8.5 or higher the degree of dissociation is assumed to be 100%:

\[
\alpha = 1 - \frac{[\text{COOH}]}{[\text{COOH}] + [\text{COO}^-]} = 1 - \frac{[\text{H}^+]_{\text{added}} - [\text{H}^+]_{\text{free}}}{[\text{acid groups}]}.
\]

Herein α is the degree of dissociation, [H⁺]₀ is the proton concentration at the beginning of the experiment (before any HCl was added), [H⁺]_{total added} is the amount of protons added, and

\[
\text{TABLE 1}
\]

<table>
<thead>
<tr>
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</tbody>
</table>
[H$^+$]$_{\text{free}}$ represents the concentration of free protons after every injection.

The main purpose of these potentiometric titration experiments was to determine the approximate degree of dissociation of the carboxylic acid groups at pH 5. From Fig. 1 it can be seen that the degree of dissociation at pH 5 is about 45 to 65%. Only a rough approximation for the degree of dissociation can be made.

**Steady-State Fluorescence Spectroscopy**

Since pyrene is a hydrophobic molecule, it partitions preferentially into an apolar environment (e.g., hydrophobic microdomains) (29, 30). The ratio of the intensity of the first peak and the third peak in the fluorescence spectrum is characteristic of the environment pyrene senses at its binding sites. In pure water this ratio is about 2, while this ratio is much lower in an apolar medium [e.g., hexane, 0.6; dodecane, 0.6; ethanol, 1.1; chloroform, 1.3 (31)]. In the hydrophobic microdomains formed by the polymers under study this ratio is circa 1.2 (32). The recorded value of $I_1/I_3$ is an average of that reported for pyrene in the microdomains and pyrene in the aqueous phase. However, since the equilibrium constant of pyrene in the aqueous phase and in micelles is of the order of $10^6$ M$^{-1}$ (29), pyrene is expected to reside almost exclusively in the microdomains.

In the case of conventional surfactants (SDS, CTAB, etc.) one can plot the $I_1/I_3$ ratio against the concentration of surfactant to determine the critical micelle concentration (CMC) as the concentration where the $I_1/I_3$ ratio drops from 2.0 to 1.2. For surfactant aggregates this should occur over a rather narrow concentration range. However, this is not the case for polymeric surfactants (16-18). Here the $I_1/I_3$ ratio decreases over a larger concentration range; therefore two arbitrary borders were chosen. The ST (start of transition) indicates the concentration where the $I_1/I_3$ ratio drops below 1.7. The concentration where the $I_1/I_3$ ratio has dropped to 1.5 is called ET (end of transition).

Some spectroscopic probes are known to induce aggregation (30). Kevelam (24) compared the CMC of an anionic surfactant as determined using steady-state fluorescence spectroscopy, isothermal titration microcalorimetry, and conductivity and found that the CMC was the same within experimental error for the three techniques.

For all polymers under various conditions (salt/no salt, low pH/high pH) the $I_1/I_3$ ratios at the highest concentration employed were typically between 1.2 and 1.3, indicating that the hydrophobic domains formed are relatively apolar (micelle-like) (29). At pH 5 in the absence of sodium citrate the concentration range for aggregation is about 2 mM. From Fig. 2 it is clear that with increasing headgroup size (i.e., increasing molecular weight) the transition concentration increases.

The almost linear dependence of the transition concentration on headgroup size is similar to that found for conventional non-ionic (polymeric) surfactants (9, 10, 13, 16, 17, 33). From the literature and our potentiometric experiments it can be derived that electrostatic interactions between polymer backbones are reduced to a large extent at pH 5 due to protonation and counterion binding (34). (Fig. 3).

At pH 9 all the carboxylic acid groups are deprotonated. However, counterion binding is expected to be about 65% (30, 33, 35). We find that with increasing headgroup size the ST increases dramatically (Fig. 4) from 6.6 mM for the polymer with $M_n$ 800 to 48 mM for the polymer with $M_n$ 1500.

![FIG. 1. Degree of dissociation ($\alpha$) versus pH at 298 K. $M_n$ 800, [polymer]$_0$ = 15.5 mM (■); $M_n$ 6700, [polymer]$_0$ = 19.9 mM (▲). At pH 5 the degree of protonation of the polymeric backbone is 45 to 65%.](image1)

![FIG. 2. Plot of CAC versus molecular weight determined by steady-state fluorescence, pH 5, 298 K, [sodium citrate] = 0 M: ST (■), ET (▲). The lines are drawn to guide the eye.](image2)

![FIG. 3. Models for the polymer coil at pH 5 for a short polymer (left) and a long polymer (right). Like surfactants with a poly (ethylene oxide) headgroup a random coil is expected.](image3)
Due to the introduction of strong electrostatic repulsion there is a large increase in the concentration range for aggregation (up to about 80 mM). However, if the polymer molecular weight is increased beyond $M_n$ 1500, this transition concentration decreases again with increasing molecular weight. This peculiar trend can be ascribed to a so-called “self-salting-out effect.” Near the hydrophobic domains the polymer chains are relatively close together, sensing a large electrostatic repulsion. This repulsion can be reduced only by an increase in counterion binding. Therefore fewer sodium ions are available for counterion binding to the end of the chains. However, the concentration of sodium counterions can be lower at the end of the polymers since the chains are further apart. The energetic strain associated with such a peculiar intrabackbone counterion distribution is less for higher-molecular-weight polymers, hence the CAC is lower. As illustrated in Fig. 5, the chains are expected to possess an extended coil formation.

At pH 9 and in the presence of 1 M sodium citrate, no distinct maximum in the ST or ET is observed (Fig. 6). The polymer of lowest molecular weight has a low ST (0.1 mM), but then the ST and ET remain more or less constant for polymers with $M_n$ between 1500 and 12,500. The ST for the polymer with $M_n$ 31,700 is somewhat lower.

By adding 1 M sodium citrate at pH 9 the transition concentration drops to about 1-2 mM, due to increased counterion binding and concomitant reduced electrostatic repulsion. However, since Fig. 6 does not exhibit a linear trend like that shown in Fig. 2 there still is some electrostatical repulsion left, and the “self-salting-out effect” is also reduced since no sharp maximum is observed at a characteristic molecular weight. This behavior can be rationalized in terms of a model that is a combination of the models suggested to operate for pH 5 and pH 9 in the absence of salt. The polymer chains will be sticking out into the solution due to electrostatic repulsion, but the coil will be less extended as that for the pH 9 solutions (Fig. 7).

**Microcalorimetry**

Isothermal titration calorimetry was employed as an additional probe-free technique to investigate the aggregation behavior of these hydrophobically modified polymers. The ST and ET obtained from the microcalorimetry experiments are labeled ST’ and ET’, respectively. In view of the fact that the details of the aggregation process at low concentration are not known we choose ST’ and ET’ rather arbitrarily (37). On injection of a small amount of a concentrated solution ($S_1$) into water ($S_2$) there is a characteristic heat effect. After every new injection of $S_1$ into

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**FIG. 4.** Plot of transition concentration versus molecular weight as determined using steady-state fluorescence spectroscopy, pH 9, 298 K, [sodium citrate] = 0 M: ST (■), ET (□). The lines are drawn to guide the eye.

**FIG. 5.** Suggested models for the polymer coil at pH 9 for a short polymer (left) and a long polymer (right). Due to large electrostatic repulsion between the negatively charged acid groups the polymeric backbones are far from each other.

**FIG. 6.** Plot of CAC versus molecular weight determined using steady-state fluorescence spectroscopy, pH 9, 298 K, [sodium citrate] = 1 M: ST (■), ET (□). The lines are drawn to guide the eye.

**FIG. 7.** Suggested models for the polymer coil at pH 9 in the presence of 1 M sodium citrate for a short polymer (left) and a large polymer (right). Electrostatic repulsion is largely reduced, but still present.
FIG. 8. Typical microcalorimetric titration experiment. $M_n$ 3000, pH 9, 298 K, 1 M sodium citrate. Micellization is endothermic.

$S_2$ the heat effect is smaller since the concentration difference between $S_1$ and $S_2$ is smaller. This is the heat effect of dilution. In the case of $S_1$ being a concentrated surfactant solution there are two separate heat effects. Initially the micelles from the solution in the syringe break up, giving an additional heat effect. However, when the CAC is reached the micelles no longer break up completely and the heat effect of micellization diminishes, while the heat effect of dilution persists throughout the whole experiment. The data in Fig. 8 reveal that aggregation of hydrophobically modified poly(sodium acrylate)s is not highly cooperative (shown by the gradual disappearance of $\Delta_{\text{demicellization}} H$), so that aggregates grow between the ST’ and ET’.

Microcalorimetry experiments with non-hydrophobically modified poly (sodium acrylate) have been performed to establish solely the enthalpic effects of polymer dilution. Straight lines for $\Delta H_{\text{obs}}$ as a function of polymer concentration are observed, providing additional evidence that the break in Fig. 8 obtained for hydrophobically modified polymers is due to (de)micellization (Fig. 9). The magnitude of the heat effects shown in Fig. 9 is dependent on the concentration of the stock solution. This effect comes from the closer packing and hence the stronger interaction between solutes in a more concentrated solution (38).

In the case of the polymers with $M_n$ 31,700 and 800 no well-defined ST’ and ET’ values were observed. For the short polymer the heat effect per injection was too small. For the longer polymer no break could be observed, as a result of relatively large heats of dilution compared with the heat effect of demicellization. For the other four polymers the ST’ and ET’ could be determined easily. We conclude that the results of the steady-state fluorescence and microcalorimetry experiments are consistent with each other. The transition concentrations are of the same order of magnitude (Fig. 10) at both pH 5 and pH 9 (Table 2).

Figure 11 shows a typical dilution experiment at 50°C. Figure 12 shows the ST’ and ET’ values as a function of temperature. Throughout the temperature range, ST’ and ET’ remain about constant, except for the ET’ at 30 and 35°C. This independence of the critical aggregation concentration on temperature has been observed for other surfactants as well (33, 39).

The heats of dilution and demicellization increase almost linearly with temperature, even becoming endothermic at 45°C (Fig. 13). At 25°C micellization is endothermic, showing that micellization is entropy driven, while at higher temperatures

![Graph](image1)

![Graph](image2)

**TABLE 2**

Comparison of Transition Concentrations as Determined Using Two Different Techniques

<table>
<thead>
<tr>
<th></th>
<th>pH 5, no salt, $M_n$ 3000</th>
<th>pH 9, no salt, $M_n$ 800</th>
</tr>
</thead>
<tbody>
<tr>
<td>ST$\text{fluorescence}$</td>
<td>0.4</td>
<td>6.6</td>
</tr>
<tr>
<td>ET$\text{fluorescence}$</td>
<td>1.2</td>
<td>12.5</td>
</tr>
<tr>
<td>ST$\text{microcalorimetry}$</td>
<td>0.5</td>
<td>5.0</td>
</tr>
<tr>
<td>ET$\text{microcalorimetry}$</td>
<td>1.1</td>
<td>22.7</td>
</tr>
</tbody>
</table>
SOLUTION BEHAVIOR OF MODIFIED POLY(SODIUM ACRYLATE)s

FIG. 11. Typical plot for the titration of hydrophobically modified poly(sodium acrylate)s into a 1 M sodium citrate solution at high temperatures (here 50°C), showing that $\Delta H_{\text{mic}}$ becomes exothermic with increasing temperature.

FIG. 12. Plot of CAC as a function of temperature: ST' ($\square$), ET' ($\bigcirc$). The lines are drawn to guide the eye.

FIG. 13. Plots of $\Delta H$ versus temperature. $M_n = 3000$, pH 9, 1 M sodium citrate: $\Delta H_{\text{dilution}}$ ($\blacksquare$), $\Delta H_{\text{dilution}} + \Delta H_{\text{demicellization}}$ ($\Box$), $\Delta H_{\text{micellization}}$ ($\bullet$).

micellization is exothermic, implying that micellization is enthalpy driven. The (almost linear) decrease in the enthalpy of micellization with temperature has been observed previously for ionic surfactants (33). From the dependence of the enthalpy of micellization on temperature it is clear that the heat capacities are negative as usually observed (38).

CONCLUSIONS

Hydrophobically mono-endcapped poly(sodium acrylate)s form hydrophobic microdomains in water. This is concluded on the basis of both steady-state fluorescence spectroscopic and isothermal titration microcalorimetric studies on poly(sodium acrylate)s with a linear C$_{12}$-alkyl chain attached specifically at the end of the polymer. However, there is no well-defined CMC, but rather a gradual transition from a micelle-free solution to a micellar solution. Steady-state fluorescence spectroscopy indicates that the microdomains are rather hydrophobic: the micropolarity at the pyrene binding sites is ethanol-like. At pH 5 in the absence of salt and at pH 9 in the presence of 1 M sodium citrate the CAC is in the range 0.2 to 2.4 mM. However, at pH 5 there is a linear increase in the transition concentration with headgroup size (i.e., molecular weight), due to an increase in steric and electrostatic repulsion between polymer main chains. At pH 9, in the absence of salt, the transition concentration is in the range 1 to 80 mM, with a sharp maximum in the transition concentration for the polymer of $M_n = 1500$. This high transition concentration is caused by a large electrostatic repulsion between carboxylate groups, the fully ionized polymers have an extended coil conformation. For the larger polymers there is a so-called “self-salting-out effect,” which consists of a concentration gradient of sodium counterions toward the hydrophobic domain, thereby increasing the sodium concentration near the hydrophobic core/hydrophilic corona interface and increasing the counterion binding and reducing the electrostatic repulsion at this interface. This effect is larger for the larger polymers because of the higher total sodium concentration and the less steep counterion concentration gradient. When polymer chains are in close proximity counterion binding is increased, whereas when polymer chains are far apart the counterion binding is reduced.

When salt is added the total counterion binding is increased. This lowers the electrostatic repulsion between the chains and reduces the transition concentration. The “self-salting-out effect” is also decreased, broadening the maximum found for the salt-free solutions and rendering the maximum less pronounced.

ACKNOWLEDGMENTS

We thank Wilma du Chatinier (Unilever Research Vlaardingen) for performing the GPC measurements. Anno Wagenaar is acknowledged for his help in performing the NMR experiments.

REFERENCES

22. Determined by viscometry using an Ubbelohde capillary viscometer, immersed in a water bath with well-controlled temperature. The capillary tube was connected to a Schott AVS 400 automatic measuring system.