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A New Method for the Synthesis of Nonsymmetric Dinucleating Ligands by Aminomethylation of Phenols and Salicylaldehydes

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Monoaminomethylated phenols 5-7 and symmetrically diaminomethylated phenols 8 and 9 were prepared in a one-step procedure from p-cresol, formaldehyde, and a variety of secondary amines by making use of the aromatic Mannich reaction. Nonsymmetric diaminomethylated phenols 10 and 11 were prepared by a sequential aromatic Mannich reaction using p-cresol, formaldehyde, and two different secondary amines. Alternatively, nonsymmetric diaminomethylated phenols 20-24 were prepared by aminomethylation of 5-methylsalicylaldehyde followed by (a) condensation with a primary amine and subsequent reduction or (b) reductive amination with a secondary amine. Monoaminomethylated and (non)symmetric diaminomethylated phenols are excellent ligands for the synthesis of mono- and dinuclear transition metal complexes as is illustrated by the isolation of mononuclear iron(III) complex 25 and nonsymmetric dinuclear copper(II) complex 26.

Introduction

Aminomethylated phenols are becoming increasingly important as ligands for both mononuclear as well as dinuclear transition metal complexes. Mononuclear iron complexes with monoaminomethylated phenols as ligands have proven to be excellent mimics for the active site of iron-tyrosinate protein. On the other hand, diamino- and tetramino-substituted phenols would force the two metal ions in the resulting complex into coordination environments. This will ultimately result in the formation of more realistic enzyme mimics.

In this paper we present new methodology for the synthesis of nonsymmetric diaminomethylated phenols via the aromatic Mannich reaction. Moreover we will show that a variety of mononucleating and symmetrically dinucleating ligands are obtained, via this aromatic Mannich reaction, in a one-step procedure from commercially available phenols and synthetically easily accessible secondary amines.

The synthetic routes to diaminomethylated phenols usually involve bishydroxymethylation of a phenol, subsequent substitution of the benzyl hydroxyl group for a chlorine using either SOCl2 or HCl and finally substitution of the benzylic chloride by an amine yielding the diaminomethylated phenol (Scheme 1, route a).

The sequential aromatic Mannich reaction is a major improvement of this three-step reaction sequence (Scheme 1, route b). Starting from a phenol, both mono- and dinaminomethylated phenols are accessible. In the aromatic Mannich reaction a phenol, a secondary amine, and formaldehyde are reacted, usually in an alcohol or an alcohol/water mixture, to provide a mononuclear complex of the monophenol in a single step. Formaldehyde can be regarded as a one-carbon donor connecting the phenol and the


secondary amine. We will show that a sequential aromatic Mannich reaction, using two different secondary amines, provides nonsymmetric dinucleating ligands. In this way ligands are created that either have a different number of amine groups or amine groups which are chemically distinct on both sides of the bridging phenol moiety.

An alternative approach comprises the synthesis of monoaminomethylated salicylaldehydes. The aldehyde functional group in these molecules offers two possible pathways for transformation into nonsymmetric dinucleating ligands, i.e., either condensation with a primary amine, followed by reduction, or reductive amination with a secondary amine.

Results and Discussion

Monoaminomethylated and Symmetrically Diaminomethylated Phenols. In order to examine the feasibility to use di- and trimines with two different amine functionalities in the aromatic Mannich reaction, the monoaminomethylation of p-cresol was studied. Monoaminomethylated phenols 5-7 (Scheme 2) were prepared by refluxing a mixture of p-cresol, the appropriate secondary amine, and paraformaldehyde or formalin solution in MeOH or i-PrOH. Although the p-cresol/secondary amine/formaldehyde ratio commonly was 1/1/1, in the case of 7 the yield was improved by using an excess of N,N-bis(2-pyridylmethyl)amine (2 equiv with respect to p-cresol). In the case of 7 a water-rich solvent mixture was also used (EtOH/H2O, 5/12) instead of pure MeOH or i-PrOH. The use of water-rich media is known to increase the rate of the aromatic Mannich reaction.15

The reactions can easily be monitored by TLC using CH2C12/triethylamine or EtOAc/hexane mixtures as eluent. Along with the formation of the monoaminomethylated p-cresol, the formation of small amounts of diaminomethylated p-cresol was observed. The monoaminomethylated p-cresols 5-7 can be isolated relatively easily from the reaction mixture containing small amounts of unreacted p-cresol, secondary amine, and diaminomethylated p-cresol. The methanol is removed in vacuo and the residue is purified by using column chromatography on silica (CH2C12/MeOH/triethylamine or EtOAc/hexane).

Monoaminomethylated phenols 5 and 6 are low-melting solids whereas 7 exists as an oil (see Experimental Section).

The attachment of a secondary amine to p-cresol via a methylenic carbon atom, which originates from formaldehyde, resulting in the formation of monoaminomethylated p-cresols 5-7, can easily be proven from NMR spectral data. Both the 1H NMR as well as the 13C NMR spectra are very indicative of the presence of the methylenic CH2 group in the aminomethylated p-cresol. The methylenic H-atoms of 5-7 all give a single resonance in the 3.6-3.8 ppm region. In the 13C NMR spectra the methylenic carbon resonances of 5-7 are in the 56-63 ppm region.

Symmetrically diaminomethylated phenols 8 and 9,14b with two diamine ligating groups and two triamine units respectively (Scheme 3), were obtained by refluxing 1 equiv of p-cresol, 3.6 equiv of secondary amine, and 3.6 equiv of formaldehyde (either paraformaldehyde or formaldehyde in water) in ethanol/water mixtures. The reactions could easily be monitored by TLC. Compound 8 was readily purified by column chromatography and was obtained as a viscous oil in excellent yield. Compound 9 was also obtained as a viscous oil after chromatographic separation although some difficulties were encountered in separating 9 from N,N-bis(2-pyridylmethyl)amine (see Experimental Section).

Nonsymmetric Diaminomethylated Phenols by a Sequential Aromatic Mannich Reaction. Having demonstrated that both monoaminomethylated and symmetrically diaminomethylated phenols with two or three amine-containing chelating units are readily accessible, we turned to the synthesis of nonsymmetric dinucleating ligands.

Starting from monoaminomethylated phenol 6 with a free ortho-position, the synthesis of nonsymmetric diaminomethylated phenols via an additional Mannich reaction was examined (Scheme 4).

The Mannich reaction with 6 was performed with pyrrolidine, N-methylpiperezine, and N,N-bis(2-pyridyl-
methyl)amine. Nonsymmetric diaminomethylated phenols 10 and 11 could indeed be prepared via this route although isolated in rather moderate yields. Both pure 10 and 11 were obtained as viscous oils after chromatographic separation but attempts to crystallize these compounds from ether/hexane mixtures were unsuccessful.

Both the \(^1\)H NMR as well as the \(^1\)H NMR spectra are very informative with regard to the nonsymmetric nature of these ligands and are very distinct from the NMR spectra of their symmetric analogues. In the \(^1\)H NMR spectra two distinct resonances for the \(H^3\) and \(H^5\) aromatic protons were observed. Moreover different resonances for the two sets of protons at the benzylic positions in compounds 10 and 11 are observed. In the \(^1\)C NMR spectra the resonances of the various benzylic and the phenyl carbon atoms are also well separated.

Unfortunately, aminomethylation of 6 with \(N,N\)-bis-(2-pyridylmethyl)amine did not yield the desired nonsymmetric ligand 12. The only product isolated was monoaminomethylated phenol 7. An explanation for this behavior might be the fact that the aromatic Mannich reaction is a reversible reaction,\(^{10}\) leading to an exchange of \(N\)-methyl-\(N\)-(2-pyridylmethyl)amine for \(N\)-methyl-N-(2-pyridylmethyl)amine in 6. This exchange reaction of secondary amines could also be responsible for the rather moderate yields of 10 and 11.

**Aminomethylation of Salicylaldehydes.** In order to improve the flexibility (and yields) in particular with respect to nonsymmetric diaminomethylated phenols containing triamine units, an alternative route, involving the Mannich reaction on salicylaldehydes, was studied. Recently it was shown that aminomethylation of 5-bromo salicylaldehyde was possible without affecting the aldehyde functional group.\(^{16}\)

Encouraged by this finding we expected that aminomethylated salicylaldehydes could easily be condensed with a variety of primary amines affording aminomethylated arylimines which are nonsymmetric dinucleating ligands.

In initial experiments it was shown that aminomethylation of salicylaldehyde 13 with two equivalents of 1-methylpiperazine afforded 3,5-diaminomethylated salicylaldehyde 14 (Scheme 5). When 1 equiv of 1-methylpiperazine was used, mixtures of 3- and 5-aminomethylated salicylaldehyde 15a and 15b were obtained (Scheme 5). Although the desired 3-aminomethylated salicylaldehyde 15a could be separated from the 5-aminomethylated salicylaldehyde 15b by column chromatography, a more selective reaction was required.

To prevent the undesired aminomethylation at the 5-position, an obvious choice is the use of 5-methylsalicylaldehyde (16). Aminomethylation of 16 with diamines and triamines is a very facile reaction affording 3-aminomethylated 5-methylsalicylaldehydes in good yields. The aminomethylation of 16 with 1-methylpiperazine to yield 18 and the aminomethylation of 16 with bis[2-(pyrrolidin-1-yl)ethyl]amine (17)\(^{17}\) to afford 19 illustrate this new route (Scheme 6). Both 18 and 19 are precursors for nonsymmetric dinucleating ligands. All monoaminomethylated salicylaldehydes (15a, 15b, 18, and 19) are solids and can be crystallized from ether/hexane mixtures.

The conversion of 18 and 19 into nonsymmetric dinucleating ligands can be accomplished in two ways: (a) the aminomethylated salicylaldehydes can be condensed with primary amines affording aminomethylated imines. Subsequently these imines can be reduced to the corresponding amines; (b) alternatively, the aminomethylated salicylaldehydes can be converted to a nonsymmetric dinucleating ligand by reductive amination with a secondary amine.

**Condensation of Aminomethylated Salicylaldehydes with Primary Amines.** Condensations of 18 (Scheme 7) with the primary amines 2-(aminomethyl)pyridine and bis(2-pyridyl)methylamine proceed readily when the reactions are performed in MeOH yielding the corresponding imines. Upon addition of the primary amine to the aminomethylated aldehyde, an intense yellow color is developed. The conversions can be monitored by TLC and are usually completed within 1 h. It is not necessary to isolate the imines prior to conversion into amines 20 and 21. Hydrogenation using \(H_2\), Pd/C or reduction with NaBH\(_4\) followed by chromatographic purification affords the corresponding amines 20 and 21.

was in accordance with the stoichiometry \([\text{Fe}(L_1)_{2}]^{2+}\) [C1041

The flexibility toward a variety of multidentate ligands is illustrated in Scheme 7. In a reverse sequence starting with monoaminomethylated salicylaldehyde 18, containing a triamine unit, a diamine unit can be attached. Thus condensation of 19 with 2-(aminomethyl)pyridine followed by NaBH₄ reduction of the in situ formed imine affords nonsymmetric dinucleating ligand 22.

Reductive Amination of Aminomethylated Salicylaldehydes. A second route to nonsymmetric diamino- or trimethylated phenols involves reductive amination of 3-aminomethylated salicylaldehydes. A typical example is given in Scheme 8. Aminomethylated salicylaldehyde 18 could be reductively aminated with pyrrolidine in MeOH using NaCNBH₃ as the reducing agent affording the nonsymmetric dinucleating ligand 23. This one-pot procedure allows rapid access to multidentate ligands although yields of this route need to be optimized and the scope with respect to a variety of mono-, di-, and triamines will be examined in due course.

Preparation of a Mononuclear Iron(II) and a Nonsymmetric Dinuclear Copper(II) Complex. As expected several of these new compounds are excellent multidentate ligands for the binding of transition metal ions. Mononuclear iron(III) complex 25 was synthesized by reaction of 2 equiv of aminomethylated phenol 6 with 1 equiv of Fe(ClO₄)₂·10 H₂O in methanol in the presence of triethylamine (Scheme 9). Elemental analysis was in accordance with the stoichiometry \([\text{Fe}(L_1)_{2}]^{2+}\) [ClO₄].

A sequential aromatic Mannich reaction on p-cresol with two different secondary amines allows the formation of nonsymmetric dinucleating ligands 10 and 11 in moderate yields. These moderate yields are probably caused by the reversibility of the aromatic Mannich reaction. Despite this drawback the route allows rapid access to multidentate ligands to study coordination behavior.

High yield and synthetically very facile alternative was developed along two lines. Aminomethylation of 5-methylsalicylaldehyde (16) with the secondary amines 1-methylpiperazine and bis[2-(pyrrolidin-1-yl)ethyl]amine yields 18 and 19 being precursors for nonsymmetric dinucleating ligands. These can be converted to nonsymmetric dinucleating ligands 20–22 by condensation with primary amines followed by reduction or by reductive amination with secondary amines such as pyrrolidine to yield 23.

The ligating capability of these new nonsymmetric dinucleating ligands is illustrated by the isolation of iron(III) complex 25 and dicopper(II) complex 26.

Concluding Remarks

The aromatic Mannich reaction has proven to be a very useful tool in the synthesis of monoaminomethylated phenols 5–7 (mononucleating ligands) and symmetrically diamino- or trimethylated phenols 8 and 9 (dinucleating ligands). The aromatic Mannich route is a very short and straightforward approach compared to the usual synthetic strategies for aminomethylated phenols. Yields are good and the reaction is attractive from the point of view of flexibility and atom economy as only water is formed as a byproduct circumventing halogenated intermediates.

Experimental Section

General. Melting points are uncorrected. 1H NMR spectra were obtained at 200 MHz. 13C NMR spectra were obtained at 50.92 MHz. The NMR spectra were recorded in CDCl₃ unless stated otherwise. Elemental analyses were performed in the Microanalytical Department of this laboratory. Mass spectra (HRMS) were obtained on an AEI-MS-902 mass spectrometer.
MeOH and EtOH were distilled from magnesium and stored over 3-A sieves. CH₂Cl₂, ether, and hexane were distilled from P₂O₅ and stored over 4-A sieves. Tritylamine was stored over solid KOH. All other reagents and solvents were of commercial grade and used without further purification. N-Methyl-N-(2-pyridylmethyl)amine,¹⁸ N-N-bis-(2-pyridylmethyl)amine,²⁴ bis-(2-pyridylmethyl)amine,²⁵ 2-pyridylmethylaldehyde,²⁶ and dimethyl iminodiacetate hydrochloride were prepared according to literature procedures.

4-Methyl-2-[[4-methylpiperazin-1-yl]methyl]phenol (5). To 1-methylpiperazine (1.01 g, 10.1 mmol) was added paraformaldehyde (0.30 g, 10.0 mmol). After the mixture was stirred for 1 h at 80 °C, p-cresol (1.05 g, 10.0 mmol) dissolved in methanol (10 mL) was added. The mixture was refluxed for 20 h. The methanol was evaporated to give a yellow oil which was purified by column chromatography on silica gel (CH₂Cl₂/MeOH/triethylamine, 50/5/1) to afford 5 (1.63 g, 74% yield) as a colorless oil. Crystallization from n-hexane afforded 5 as white crystals: mp 61.3–62.3 °C; ¹H NMR (200 MHz) δ 2.23 (s, 3H), 2.29 (s, 3H), 2.54 (br, 8H), 2.66 (s, 2H), 6.71 (d, 1H, J = 8.1 Hz), 6.78 (br, s, 1H), 6.96 (dd, 1H, J = 3.1, 8.1 Hz); ¹³C NMR (50 MHz) δ 20.4 (q), 24.9 (s), 54.9 (t), 63.1 (t), 115.7 (d), 120.8 (d), 128.1 (s), 129.2 (d), 156.3 (s); HRMS calcd for C₉H₉NO₂ 220.158, found 220.158. Anal. Calc'd for C₉H₉NO₂: C, 70.87; H, 9.15; N, 12.72. Found: C, 70.97; H, 9.11; N, 12.42.

4-Methyl-2-[[(N-methyl-N-(2-pyridylmethyl)amino)methyl]phenol (6). Following the procedure for 5, reaction of N-methyl-N-(2-pyridylmethyl)amine (2.40 g, 13.7 mmol), paraformaldehyde (0.30 g, 10.0 mmol), and p-cresol (2.18 g, 10.0 mmol) in i-PrOH (10 mL) after a reaction time of 2 d, workup, and chromatography on silica gel (ethyl acetate/hexane, 1/1) afforded 6 (2.54 g, 55% yield) as a slightly yellow oil. Crystallization from ether/hexane afforded 6 as white crystals: mp 52.9–53.5 °C; ¹H NMR (200 MHz) δ 2.25 (s, 3H), 3.23 (s, 3H), 3.74 (s, 2H), 3.79 (s, 2H), 6.77 (d, 1H, J = 8.1 Hz), 8.63 (s, 1H), 6.99 (dd, 1H, J = 8.1, 1.7 Hz), 7.21 (m, 1H), 7.36 (d, 1H, J = 8.1 Hz), 7.70 (m, 1H), 8.59 (m, 1H), 10.80 (br, 1H); ¹³C NMR (50 MHz) δ 20.4 (q), 41.7 (q), 60.7 (d), 62.8 (t), 115.9 (d), 121.7 (s), 122.4 (d), 125.2 (d), 129.1 (s), 129.2 (d), 129.4 (d), 149.2 (s), 157.4 (s); HRMS calcd for C₁₇H₁₉NO₃ 342.142, found 342.142. Anal. Calc'd for C₁₇H₁₉NO₃: C, 74.35; H, 7.49; N, 11.65. Found: C, 74.34; H, 7.56; N, 11.45.

2-[([N-Bis(2-pyridylmethyl)amino)methyl]-4-methylphenol (7). To N,N-bis(2-pyridylmethyl)amine (1.75 g, 8.8 mmol) in EtOH/H₂O (5 mL/12 mL) was added p-cresol (0.486 g, 4.5 mmol) and formalin solution (0.7 mL, 37% in H₂O). The two-phase reaction mixture was refluxed for 3 d and allowed to cool to rt. The reaction mixture was partitioned between CH₂Cl₂ (100 mL) and H₂O (50 mL), and the H₂O layer was extracted with CH₂Cl₂ (2 × 20 mL). The combined organic layers were washed with Na₂SO₄ and evaporated to leave a crude oil. Column chromatography on silica gel (CH₂Cl₂/MeOH/triethylamine, 95/5/1) afforded 7 (0.954 g, 66% yield) as a slightly yellow oil: ¹H NMR (200 MHz) δ 2.25 (s, 3H), 3.75 (s, 2H), 3.86 (s, 4H), 6.82 (d, 1H, J = 8.1 Hz), 6.87 (br, s, 1H), 6.98 (dd, 1H, J = 8.1, 1.8 Hz), 7.14 (m, 2H), 7.34 (d, 2H, J = 7.7 Hz), 7.61 (dd, 2H, J = 7.7, 1.8 Hz), 8.55 (m, 2H); ¹³C NMR (50 MHz) δ 20.4 (q), 57.0 (t), 59.1 (t), 116.2 (d), 122.2 (d), 122.5 (d), 127.6 (d), 129.5 (d), 130.7 (d), 131.8 (d), 146.8 (s), 151.1 (s), 156.3 (s); HRMS calcd for C₂₁H₁₈NO₄ 319.168, found 319.168.

2.5-Bis[[N-(N-methyl-N-(2-pyridylmethyl)amino)methyl]4-methylphenol (8). Following the procedure for 7, reaction of p-cresol (0.119 g, 1.10 mmol), N,N-bis(2-pyridylmethyl)amine (0.503 g, 4.12 mmol), and paraformaldehyde (0.126 g, 4.20 mmol) in H₂O/EtOH (5 mL/2 mL) after a reaction time of 3.5 d, workup, and chromatography on silica gel (CH₂Cl₂/MeOH,

and salicylaldehyde 18 (2.45 g, 20.1 mmol) in methanol (10 mL) after a reaction time of 30 h and column chromatography on silica gel (CH$_2$Cl$_2$/MeOH, 9/1) afforded a fraction of pure 15a (1.21 g, 26% yield), a fraction (0.664 g, 14% yield) consisting of a mixture of 15a and 15b (15a/15b, 7:1,1) and a fraction (0.56 g, 12% yield) consisting of a mixture of 18a and 18b (18a/18b, 1:34). All fractions were solid. Recrystallization from hexane/
ether gave 15a as slightly yellow crystals: mp 86–87.0 °C; *H NMR (200 MHz) $\delta$ 2.32 (s, 3 H), 2.49–2.61 (br, 8 H), 3.47 (s, 2 H), 6.84 (d, $J = 7.7$, 7.7 Hz, 1 H), 7.26 (m, 1 H), 7.60 (d, $J = 7.7$, 7.7 Hz, 1 H), 3.60 (q, 4 H); 13C NMR (50 MHz) $\delta$ 20.2 (q), 45.9 (q), 52.8 (t), 55.0 (t), 61.8 (t), 7.7, 7.7 Hz, 1 H), 7.26 (d, $J = 7.7$, 7.7 Hz, 1 H), 6.90 (s, 1 H), 7.11 (m, 1 H), 7.32 (d, $J = 7.7$, 7.7 Hz, 1 H), 7.61 (dd, $J = 7.7$, 7.7, 1.7 Hz, 1 H), 8.50 (m, 1 H); 13C NMR (50 MHz) $\delta$ 20.4 (q), 45.8 (q), 49.2 (t), 54.3 (t), 54.8 (t), 60.6 (t), 62.9 (t), 120 (a), 121.9 (d), 122.3 (d), 125.3 (s), 127.7 (s), 128.6 (d), 129.5 (d), 138.5 (d), 149.1 (d), 156.3 (s), 158.4 (s); HRMS calcd for C$_9$H$_{12}$N$_3$O$_2$ 191.262, found: 191.262.

4-Methyl-2-[(4-methylpiperazin-1-yl)methyl]-6-[[N-(bis-(2-pyridyl)ethyl)amino]methyl]phenol (21). To 18 (0.502 g, 2.44 mmol) in methanol (10 mL) was added bis(2-pyridyl)methylamine (0.452 g, 2.44 mmol). After stirring for 3 h, H$_2$O (15 mL) was added carefully in portions. The mixture was stirred for 3 h, HzO (15 mL) was added carefully in portions. The mixture was stirred for 1 h, dried over Na$_2$SO$_4$ and evaporated to leave a dark yellow oil. The oil was purified by chromatography on silica gel (CH$_2$Cl$_2$/MeOH, 9/1) to afford 21 (0.647 g, 63% yield) as a colorless oil: *H NMR (200 MHz) $\delta$ 2.08 (m, 4 H), 2.19 (t, 2 H), 2.57 (br, 8 H), 3.70 (e, 2 H), 3.82 (e, 2 H), 4.32 (e, 2 H), 4.38 (e, 2 H) (200 MHz) $\delta$ 20.4 (q), 45.6 (q), 47.8 (t), 52.1 (t), 54.3 (t), 54.8 (t), 60.6 (t), 62.9 (t), 121.9 (d), 122.3 (d), 126.5 (s), 127.7 (s), 128.7 (d), 129.5 (s), 136.6 (d), 149.3 (d), 156.7 (s), 158.2 (s); HRMS calcd for C$_{15}$H$_{18}$N$_3$O$_2$ 258.228, found: 258.228.

2-[(Bis[2-(pyrrolidin-1-yl)ethyl]amino)methyl]-6-[[N-(2-pyrroldin-1-yl)ethyl]amino]methyl]phenol (22). To 18 (0.159 g, 0.433 mmol) in methanol (5 mL) was added 2-(amino- methyl)pyridine (0.105 g, 1.04 mmol). After stirring for 1 h at rt, NaBH$_4$ (0.493 g, 1.30 mmol) was added in small portions. After stirring for 2 h at rt the reaction mixture was acidified to pH 1–2 using a 2 N HCl solution and stirred for another 0.5 h. After the solution was brought to pH 7–8 using a 2 N KOH solution, the mixture was extracted with CH$_2$Cl$_2$ (3 × 50 mL). The organic layers were dried over Na$_2$SO$_4$ and evaporated to leave a dark yellow oil. The oil was purified by chromatography on silica gel (CH$_2$Cl$_2$/MeOH, 9/1) to afford 22 (0.302 g, 61% yield) as a dark red oil: *H NMR (200 MHz) $\delta$ 1.76 (m, 8 H), 1.87 (m, 8 H), 2.18 (e, 2 H), 2.30 (m, 8 H), 2.57 (br, 8 H), 2.75 (m, 4 H), 2.76 (m, 4 H); 13C NMR (50 MHz) $\delta$ 23.4 (t), 48.6 (t), 54.2 (t), 55.9 (t) (200 MHz) $\delta$ 2.08 (m, 4 H), 2.19 (t, 2 H), 2.57 (br, 8 H), 3.70 (e, 2 H), 3.82 (e, 2 H), 4.32 (e, 2 H), 4.38 (e, 2 H) (200 MHz) $\delta$ 20.4 (q), 45.6 (q), 47.8 (t), 52.1 (t), 54.3 (t), 54.8 (t), 60.6 (t), 62.9 (t), 121.9 (d), 122.3 (d), 126.5 (s), 127.7 (s), 128.7 (d), 129.5 (s), 136.6 (d), 149.3 (d), 156.7 (s), 158.2 (s); HRMS calcd for C$_{15}$H$_{18}$N$_3$O$_2$ 258.228, found: 258.228.

2-[Bis[2-(pyrrolidin-1-yl)ethyl]amino]methyl]-6-[N-(2-pyrrolidin-1-yl)methyl]phenol (23). To 18 (0.246 g, 0.992 mmol) in methanol (5 mL) was added pyridilone (0.077 g, 1.09 mmol), NaOAc (0.165 g, 2.01 mmol), and NaBH$_4$/CN (0.13 g, 2.06 mmol). The reaction mixture was stirred for 18 h at rt. Using the workup procedure for 21 followed by chromatography on silica gel (CH$_2$Cl$_2$/MeOH/triethylamine, 75/25/1) to afford 23 (0.015 g, 32% yield) as a dark purple crystals. Anal. Calcd for C$_{15}$H$_{18}$N$_3$O$_2$: C, 56.5; H, 5.37; N, 8.78. Found C, 56.4; H, 5.41; N, 8.89.
Synthesis of Nonsymmetric Dinucleating Ligands

$[\text{Cu}(L')(\text{OAc})_2][\text{ClO}_4]$ (26). To 15a (0.236 g, 1.01 mmol) in MeOH (50 mL) was added 2-(aminoethyl)pyridine (0.123 g, 1.01 mmol). After stirring for 1 h at rt Cu(OAc)$_2$H$_2$O (0.40 g, 2.00 mmol) was added. After the mixture was refluxed for 1.5 h, NaClO$_4$·H$_2$O (0.200 g, 1.42 mmol) was added and the mixture was refluxed for another 0.5 h. The reaction mixture was filtered and the MeOH was removed in vacuo. A small amount of EtOH was added to the residue. The mixture was boiled for 5 min and filtered to afford 26 (0.43 g, 62% yield). Slow crystallization from CH$_2$Cl$_2$/EtOH afforded dark green crystals. Anal. Calcd for C$_{32}$H$_{27}$ClCu$_2$N$_4$O$_4$: C, 42.26; H, 4.58; Cu, 18.63, N, 8.21. Found: C, 41.47; H, 4.66; Cu, 18.81, N, 8.10.

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Supplementary Material Available: Copies of $^1$H NMR spectra of all new compounds (15 pages). This material is contained in libraries on microfiche, immediately follows this article in the microfilm version of the journal, and can be ordered from the ACS; see any current masthead page for ordering information.